

$$C = \frac{N}{V}$$

$$N = \frac{m}{M}$$

$$C_1 V_1 = C_2 V_2$$

IMPORTANT FORMULA: CONCENTRATION

$$C = \frac{N}{V} \quad \frac{\text{mols}}{L}$$

$$V_1 C_1 = V_2 C_2$$

P1

P1

Stoichiometry

Write the information under the balanced equation
use $N = \frac{m}{M}$ to find mols of the Known.

When N is found use $N = \frac{\text{unknown Coeff.}}{\text{Known Coeff.}}$

use the new N value to find mass of the unknown with $N = \frac{m}{M} \rightarrow$ what you find.

P1

P1

$\frac{W}{M}$

$$n = \frac{m}{M}$$

$$n = \frac{V}{V}$$

$$PV = nRT$$

$$\frac{P_1 V_1}{P_2 V_2} = \frac{n_1 R T_1}{n_2 R T_2}$$

Gases

Scaled Flask Calculations

$$\text{mol of compound, } n = \frac{m}{M}$$

mol of —, mol of compound \times mol of element
mass of —, mol of \times mol of e, \times mass of e
of atoms of —, mol of \times mol of $\times 6.02 \times 10^{23}$
total atoms, mol of \times # of e $\times 6.02 \times 10^{23}$

P1